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Stoichiometry problem worksheet

6th, 7th, 8th, 9th, 10th, 11th, 12th, Adult Education, HomeschoolPage 25th, 6th, 7th, 8th, 9th, 10th, 11th, 12, Adult Education, Homeschool Balancing the following chemical reactions: a. $2\text{CO} + \text{O}_2 \rightarrow 2\text{CO}_2$ b. $2\text{KNO}_3 \rightarrow 2\text{KNO}_2 + \text{O}_2$ c. $2\text{O}_3 \rightarrow 3\text{O}_2$ d. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + 2\text{H}_2\text{O}$ e. $4\text{CH}_3\text{NH}_2 + 9\text{O}_2 \rightarrow 4\text{CO}_2 + 10\text{H}_2\text{O} + 2\text{N}_2$ f. $\text{Cr}(\text{OH})_3 + 3\text{HClO}_4 \rightarrow \text{Cr}(\text{ClO}_4)_3 + 3\text{H}_2\text{O}$ Write the balanced chemical equations of each reaction: a. Calcium carbide (CaC_2) reacts with water to form calcium hydroxide ($\text{Ca}(\text{OH})_2$) and acetylene gas (C_2H_2). $\text{CaC}_2 + 2\text{H}_2\text{O} \rightarrow \text{Ca}(\text{OH})_2 + \text{C}_2\text{H}_2$ b. When potassium chlorate (KClO_3) is heated, it decomposes to form KCl and oxygen (O_2). $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ c. C_6H_6 burns in the air. $2\text{C}_6\text{H}_6 + 15\text{O}_2 \rightarrow 12\text{CO}_2 + 6\text{H}_2\text{O}$ d. $\text{C}_5\text{H}_{12}\text{O}$ burns in the air. $2\text{C}_5\text{H}_{12}\text{O} + 15\text{O}_2 \rightarrow 10\text{CO}_2 + 12\text{H}_2\text{O}$ Give the following reaction: $\text{Na}_2\text{S}_2\text{O}_3 + \text{AgBr} \rightarrow \text{NaBr} + \text{Na}_3[\text{Ag}(\text{S}_2\text{O}_3)_2]$ a. How many moles of $\text{Na}_2\text{S}_2\text{O}_3$ are needed to react completely to 42.7 g AgBr ? 0.455 mol $\text{Na}_2\text{S}_2\text{O}_3$ b. What will the volume of NaBr be produced from 42.7 g AgBr ? 23.4 g NaBr From Response: $\text{B}_2\text{H}_6 + \text{O}_2 \rightarrow \text{HBO}_2 + \text{H}_2\text{O}$ a. What volume will O_2 be needed to burn 36.1 g B_2H_6 ? 125 g O_2 b. How many moles of water are produced from 19.2 g B_2H_6 ? 1.39 H_2O mol c. Calculate the volume (in kilograms) of water generated from the combustion of 1.0 gallons (3.8 L) of gasoline (C_8H_{18}). The gasoline density is 0.79 g/mL. 4.3 kg H_2O A mole of aspartame ($\text{C}_{14}\text{H}_{18}\text{N}_2\text{O}_5$) reacts with two moles of water to produce a mole of aspartic acid ($\text{C}_4\text{H}_7\text{NO}_4$), a mole of methanol (CH_3OH) and a mole of nmol . A. What is the molecular formula of phenylalanine? $\text{C}_9\text{H}_{11}\text{NO}_2$ b. What volume of phenylalanine is produced from 378 g of aspartame? 212 g phenylalanine CO_2 is used in a closed system breathing device. It removes carbon dioxide and water from the exhaled air. The reaction to remove water is: $\text{KO}_2 + \text{H}_2\text{O} \rightarrow \text{KOH}$. KOH production is used to remove carbon dioxide by the following reactions: $\text{KOH} + \text{CO}_2 \rightarrow \text{KHCO}_3$. A. Which KO_2 mass produces 235 g O_2 ? 696 g KO_2 b. What CO_2 volume can be removed by 123 g KO_2 ? 76.1 g CO_2 Balance the following chemical reactions: Hint a. $\text{CO} + \text{O}_2 \rightarrow \text{CO}_2$ b. $\text{KNO}_3 \rightarrow \text{KNO}_2 + \text{O}_2$ c. $\text{O}_3 \rightarrow \text{O}_2$ d. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$ e. $\text{CH}_3\text{NH}_2 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O} + \text{N}_2$ Hint f. $\text{Cr}(\text{OH})_3 + \text{HClO}_4 \rightarrow \text{Cr}(\text{ClO}_4)_3 + \text{H}_2\text{O}$ Write the balanced chemical equations of each reaction: a. Calcium carbide (CaC_2) reacts with water to form calcium hydroxide ($\text{Ca}(\text{OH})_2$) and acetylene gas (C_2H_2). b. When potassium chlorate (KClO_3) is heated, it decomposes to form KCl and oxygen (O_2). c. C_6H_6 burns in the air. Hint d. $\text{C}_5\text{H}_{12}\text{O}$ burns in the air. With the following reaction: $\text{Na}_2\text{S}_2\text{O}_3 + \text{AgBr} \rightarrow \text{NaBr} + \text{Na}_3[\text{Ag}(\text{S}_2\text{O}_3)_2]$ a. How many moles of $\text{Na}_2\text{S}_2\text{O}_3$ are needed to react completely to 42.7 g AgBr ? b. What will the volume of NaBr be produced from 42.7 g AgBr ? Hint From Response: $\text{B}_2\text{H}_6 + \text{O}_2 \rightarrow \text{HBO}_2 + \text{H}_2\text{O}$ a. What volume will O_2 be needed to burn 36.1 g B_2H_6 ? b. How many moles of water are produced from 19.2 g B_2H_6 ? Calculate the volume (in kilograms) of water generated from the combustion of 1.0 gallons (L) gasoline (C_8H_{18}). The gasoline density is 0.79 g/mL. A mole of aspartame ($\text{C}_{14}\text{H}_{18}\text{N}_2\text{O}_5$) reacts with two moles of water to produce a mole of aspartic acid ($\text{C}_4\text{H}_7\text{NO}_4$), a mole of methanol (CH_3OH) and a mole of nmol . A. What is the molecular formula of phenylalanine? Hint b. What volume of phenylalanine is produced from 378 g of aspartame? KO_2 is used in a closed system breathing device. It removes carbon dioxide and water from the exhaled air. The reaction to remove water is: $\text{KO}_2 + \text{H}_2\text{O} \rightarrow \text{KOH}$. KOH production is used to remove carbon dioxide by the following reactions: $\text{KOH} + \text{CO}_2 \rightarrow \text{KHCO}_3$. A. Which KO_2 mass produces 235 g O_2 ? B. What CO_2 volume can be removed by 123 g KO_2 ? Hint If you see this message, it means that we're having trouble loading external resources on our website. If you're behind a web filter, please make sure that the *.kastatic.org and *.kasandbox.org domains are unblocked. unblock.

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